

How to take H₂ excited levels into account

Case 1: $50 < T_k < 500$ K

Hydrogen molecules (H₂) exist in two isomeric forms: para-hydrogen with total nuclear spin $I = 0$ and ortho-hydrogen with $I = 1$. In the ground electronic and vibrational state¹, the rotational levels of para-H₂ have even values of the angular momentum j while the levels of ortho-H₂ have odd j values. The ground-state of ortho-H₂ ($j=1$) lies 170 K above the ground-state of para-H₂ ($j = 0$).

In cold environments where the kinetic temperature T_k is lower than ~ 50 K and in the absence of radiative pumping, only the ground-state of the two nuclear spin isomers are significantly populated so that only para-H₂ ($j = 0$) and ortho-H₂ ($j = 1$) need to be considered as collision partners. Moreover, the rate coefficients for ortho-H₂ ($j = 1$) are generally larger than those for para-H₂ ($j = 0$), by up to an order of magnitude.

At higher kinetic temperatures $T_k > 50$ K or in the presence of radiative pumping, H₂ molecules in excited levels $j = 2, 3$, etc. can become significant new colliders. It has been generally found, however, that the dominant collisional rate coefficients for H₂ ($j > 1$) differ by less than 20-30% from those for H₂ ($j = 1$) (*e.g.* Daniel et al. 2014 and references therein). This result holds for targets in their ground vibrational state and for target rotational levels or kinetic temperatures below ~ 500 K, *i.e.* below the opening of the H₂ level $j = 2$ at 510 K. In these conditions, rotation-rotation transfers between the target species and H₂ are negligible so that H₂ remains mostly in the same rotational state during the collision (*i.e.* $j \rightarrow j$).

EMAA provides de-excitation rate coefficients (in cm^3s^{-1}) due to collisions with ‘para-H₂’ and ‘ortho-H₂’ which in practice correspond to H₂ ($j = 0 \rightarrow 0$) and H₂ ($j = 1 \rightarrow 1$), respectively (except otherwise stated). In order to include excited levels of H₂, rate coefficients for *e.g.* H₂ ($j = 2 \rightarrow 2$) can be assumed to be approximately equal to those for H₂ ($j = 1 \rightarrow 1$). **Thus, in radiative transfer calculations, the volume density (in cm^{-3}) affected to the ‘para-H₂’ collider should be that of H₂ ($j = 0$) while the volume density affected to the ‘ortho-H₂’ collider should be that of the sum of all levels of H₂ with $j > 0$.** A simple solution to estimate the H₂ level populations is to further assume that these levels are thermalized at the kinetic temperature, with a fixed or thermalized ortho-to-para ratio.

References

- Daniel F., Faure A., Wiesenfeld L., Roueff E., Lis D. C., Hily-Blant P., *MNRAS* **444** 2544-2554 (2014)
 Faure A., Lique F., Wiesenfeld L., *MNRAS* **460** 2103-2109 (2016)

¹The first vibrational level of H₂ opens at 4161.2 cm^{-1} .